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73 Proprietor: ASAHI GLASS COMPANY LTD.
1-2, Marunouchi 2-chome
Chiyoda-ku
Tokyo 100 (JP)

72 Inventor: Kushitani, Hideki, c/o Asahi Glass

Company Ltd.
Chuo Kenkyusho 1150,
Hazawa-cho,
Kanagawa-ku
Yokohama-shi,
Kanagawa-ken (JP)
Inventor: Nishizawa, Manabu, c/o Asahi Glass
Company Ltd.
Chuo Kenkyusho 1150,
Hazawa-cho,
Kanagawa-ku
Yokohama-shi,
Kanagawa-ken (JP)
Inventor: Nakao, Yasumasa, c/o Asahi Glass
Company Ltd.
Chuo Kenkyusho 1150,
Hazawa-cho,
Kanagawa-ku
Yokohama-shi,
Kanagawa-ken (JP)

74 Representative: Wächtershäuser, Günter, Prof.
Dr.
Patentanwalt,
Tal 29
D-80331 München (DE)

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Description

The present invention relates to an alkali free glass which is substantially free from alkali metal oxides and capable of being formed by a float method and which is thus useful as glass substrates for various displays or photo masks.

C.A. Vol. 95, No. 8 (August 1981), p. 238, Ref.No. 666196 and JP-A-56032347 disclose glasses with strain points of 700 °C or more.

US-A-4,255,198 discloses glasses with strain points higher than 730 °C.

Heretofore, glass substrates for various displays, particularly those having a thin metal film formed on their surface, are required to have high levels of the following properties.

- (1) Since they are exposed to high temperature during the step for the formation of the thin film, they are required to have a high strain point.
- (2) If they contain an alkali metal oxide, the alkali metal ions are likely to diffuse into the thin film, whereby the film properties are likely to deteriorate. Therefore, they are required to be substantially free from alkali metal ions.
- (3) They are required to be free from internal and surfacial defects (bubbles, striae, inclusions, pits, scratch marks, etc.).
- (4) They are required to have excellent chemical durability so as to be durable during the washing step.

In recent years, liquid crystal displays having thin film transistors (TFT) using amorphous silicon (a-Si) or polycrystalline silicon (p-Si) formed on a glass substrate, are increasing. A display substrate using the latter polycrystalline silicon, is required to have a strain point of at least 650 °C.

Further, these display substrates are frequently exposed to buffered hydrofluoric acid which is prepared by adding ammonium fluoride as a buffering agent to hydrofluoric acid, for etching e.g. SiO₂ or Si₃N₄ in a process for producing semiconductors.

Conventional alkali free glass has a low strain point or when exposed to buffered hydrofluoric acid, it becomes useless in many cases, since a reaction product formed on the surface can not easily be removed.

Glass having a strain point of at least 650 °C is disclosed in Japanese Unexamined Patent Publications No. 32038/1983, No. 132536/1986 and No. 281041/1986. However, the glass composition of Japanese Unexamined Patent Publication No. 32038/1983 contains lead oxide, whereby it is necessary to provide a removal apparatus for the mixing installation, the melting installation and the polishing installation, and such environment is not preferable as a working environment. The glass composition of Japanese Unexamined Patent Publications No. 132536/1986 and No. 281041/1986 contain excess MgO and BaO, whereby the buffered hydrofluoric acid resistance is not adequate.

Japanese Unexamined Patent Publication No. 132536/1986 discloses a glass substantially free from MgO. This glass is excellent in the buffered hydrofluoric acid resistance, but the strain point is not sufficiently high, since it contains from 3 to 12 wt% of B₂O₃.

It is an object of the present invention to solve the above problems and to provide an alkali free glass which has a strain point of at least 650 °C and is excellent in the buffered hydrofluoric acid resistance and which can easily be melted and formed and is capable of being formed by a float method and has a low thermal expansion coefficient.

The present invention provides an alkali free glass having a strain point of at least 675 °C and consisting essentially of from 58 to 61 wt% of SiO₂, from 11.5 to 15 wt% of Al₂O₃, from 1 to 2.5 wt% of B₂O₃, from 10 to 14 wt% of CaO, from 11 to 15 wt% of SrO and from 0 to 2 wt% of ZnO and being essentially free from MgO, BaO and alkali metal oxides.

Now, the present invention will be described in detail with reference to the preferred embodiments.

In this specification, the viscosity η of glass is represented by the unit of poise, and the strain point is the temperature where $\log \eta = 14.5$.

Now, the reasons for limiting the compositional ranges of the respective components will be described.

If SiO₂ is less than 55 wt%, the chemical durability of the glass tends to be poor, and the thermal expansion coefficient tends to be large, such being undesirable. On the other hand, if it exceeds 65 wt%, the melting property tends to deteriorate, and the liquidus temperature tends to increase, such being undesirable. Therefore, SiO₂ is within a range of from 58 to 61 wt%.

Al₂O₃ has a function to suppress the phase separation of the glass, to reduce the thermal expansion coefficient and to increase the strain point. If it is less than 10 wt%, this effect can not be obtained, and if it exceeds 18 wt%, the melting property of the glass tends to be poor, and the buffered hydrofluoric acid resistance tends to be poor. Therefore, Al₂O₃ is within a range of from 11.5 to 15 wt%.

B_2O_3 is a component effective for preventing formation of turbidity by buffered hydrofluoric acid and for providing a low expansion coefficient and good melting property. However, if it exceeds 3 wt%, the strain point tends to be low, such being undesirable. Therefore, B_2O_3 is within a range of from 1 to 2.5 wt%.

MgO is effective for lowering the thermal expansion coefficient of the glass. However, if it exceeds 3 wt%, turbidity by buffered hydrofluoric acid or phase separation of the glass is likely to occur, such being undesirable. Therefore, MgO is not substantially present.

CaO is effective for improving the melting property and controlling the liquidus temperature. However, if it is less than 8 wt%, such an effect tends to be inadequate, and if it exceeds 15 wt%, the thermal expansion coefficient tends to be too large, and crystals of $CaO \cdot Al_2O_3 \cdot SiO_2$ tend to precipitate, such being undesirable. Therefore, CaO is within a range of from 10 to 14 wt%.

SrO provides substantially the same effect as CaO , when incorporated. If SrO is less than 8 wt%, such an effect is little, and if it exceeds 15 wt%, the thermal expansion coefficient tends to be too large. Therefore, SrO is within a range of from 11 to 15 wt%.

ZnO is effective to lower the thermal expansion coefficient of the glass. However, if it exceeds 3 wt%, turbidity by buffered hydrofluoric acid is likely to form, and it is likely to be reduced and evaporated to form a different quality layer on the surface in the float bath during its preparation, whereby forming by a float method tends to be difficult. Therefore, ZnO is within a range of from 0 to 2 wt%.

BaO is not substantially contained, since it tends to deteriorate the buffered hydrofluoric acid resistance.

The glass of the present invention may contain not more than 5 wt% in total of ZrO_2 , P_2O_5 , TiO_2 , SO_3 , As_2O_3 , Sb_2O_3 , F and Cl in order to improve the melting property, refining property and formability of the glass in addition to the above-mentioned components.

The glass of the present invention can be produced, for example, by the following process.

Starting materials of the respective components commonly used, are mixed in the desired proportions, and the mixture is continuously introduced into a melting furnace and melted under heating at a temperature of from 1,500 to 1,600°C. This molten glass is formed into a sheet having a predetermined thickness by a float method, followed by gradual cooling and cutting.

Now, the present invention will be described in further detail with reference to Examples. However, it should be understood that the present invention is by no means restricted to such specific Examples.

30 EXAMPLES and COMPARATIVE EXAMPLES

Starting materials of the respective components were mixed to have the desired proportions as identified in Table 1 (unit: wt%), and the mixture was heated and melted in a platinum crucible at a temperature of from 1,500 to 1,600°C. The molten glass was stirred with a platinum stirrer for 2 hours for homogenizing the glass during the melting. Then, the molten glass was casted into a sheet form, followed by annealing.

The results of the measurements of the physical properties of the glass samples thereby obtained are shown in Table 2. In the Tables, samples 1, 2 and 4 represent Examples of the present invention, and samples 3 and 5 to 10 represent Comparative Examples.

The buffered hydrofluoric acid resistance was determined in such a manner that a sheet glass polished to have a mirror surface was immersed in a liquid mixture of an aqueous ammonium fluoride solution and an aqueous hydrofluoric acid solution (a solution obtained by mixing a 40 wt% ammonium fluoride aqueous solution and a 51 wt% hydrofluoric acid aqueous solution in a weight ratio of 30:1) at 25°C for 20 minutes, whereupon the haze value was measured by a haze meter (manufactured by Suga Shikenki K.K.).

As is evident from Table 2, the glass in each of the Examples has a low thermal expansion coefficient at a level of from 45 to $60 \times 10^{-7}/^\circ C$ and a relatively low temperature for $\log \eta = 2.5$ which is an index for the melting property, thus indicating that it can easily be melted. Further, the relation as an index for formability between the temperature for $\log \eta = 4.0$ and the liquidus temperature is good, thus indicating that there will be no trouble of devitrification during the forming operation.

The strain point is as high as at least 650°C, thus indicating that it is adequately durable against heat treatment at a high temperature.

The buffered hydrofluoric acid resistance is also excellent, and the glass is free from formation of turbidity.

On the other hand, samples 8 and 9 representing Comparative Examples have high strain points, and they are inferior in the buffered hydrofluoric acid resistance, although the temperature for $\log \eta = 2.5$ as an index of the melting temperature, is relatively low and the melting property is considered to be good. Further, sample 10 as a Comparative Example has a low strain point and is considered to be not durable against heat treatment at a high temperature, although the buffered hydrofluoric acid resistance is good.

Table 1

Sample No.	1	2	3	4	5	6	7	8	9	10
SiO ₂	60	60	61	60	61	58	57	55	60	60
Al ₂ O ₃	12	15	12	12	12	11	15	15	17	12
B ₂ O ₃	2	2	2	1	2	3	-	2	4	5
MgO	-	-	3	-	-	2	2	6	8	-
CaO	12	11	8	13	12	12	14	6	11	7
SrO	14	12	14	14	10	14	9	-	-	1
BaO	-	-	-	-	-	-	-	6	-	15
ZnO	-	-	-	-	3	-	3	10	-	-

Sample No.s 3 and 5 to 10 represent Comparative Examples

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Table 2

Sample No.	1	2	3	4	5	6	7	8	9	10
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)	56	52	49	56	52	41	59	44	44	47
Temp. where $\log \eta = 2.5 (^{\circ}\text{C})$	1,447	1,503	1,463	1,440	1,450	1,498	1,430	1,386	1,419	1,476
Temp. where $\log \eta = 4.0 (^{\circ}\text{C})$	1,201	1,237	1,218	1,197	1,210	1,213	1,170	1,163	1,190	1,204
Liquidus temp. ($^{\circ}\text{C}$)	1,200	1,215	1,238	1,194	1,228	1,188	1,165	1,168	1,185	1,150
Strain point ($^{\circ}\text{C}$)	680	685	670	690	670	650	670	665	675	635
Buffered hydrofluoric acid resistance (%)	6	7	7	7	4	2	10	35	28	3

Sample No.s 3 and 5 to 10 represent Comparative Examples

The glass of the present invention is capable of being formed by a float method. Further, it is free from formation of turbidity by buffered hydrofluoric acid and has high heat resistance and a low thermal expansion coefficient. Thus, it is suitable for use in the applications where such properties are required, for example, as display substrates, photo mask substrates or TFT type display substrates.

Claims

1. An alkali free glass having a strain point of at least 675 °C and consisting essentially of from 58 to 61 wt% of SiO₂, from 11.5 to 15 wt% of Al₂O₃, from 1 to 2.5 wt% of B₂O₃, from 10 to 14 wt% of CaO, from 11 to 15 wt% of SrO and from 0 to 2 wt% of ZnO and being essentially free from MgO, BaO and alkali metal oxides.

Patentansprüche

1. Alkalifreies Glas, das einen unteren Kühlpunkt von mindestens 675 °C hat und das im wesentlichen aus 58 bis 61 Gew.-% SiO₂, 11,5 bis 15 Gew.-% Al₂O₃, 1 bis 2,5 Gew.-% B₂O₃, 10 bis 14 Gew.-% CaO, 11 bis 15 Gew.-% SrO und 0 bis 2 Gew.-% ZnO besteht und das im wesentlichen frei von MgO, BaO und Alkalimetalloxiden ist.

Revendications

1. Verre exempt d'alcalis présentant un point de trempe d'au moins 675 °C, et essentiellement constitué de 58 à 61% en poids de SiO₂, de 11,5 à 15% en poids de Al₂O₃, de 1 à 2,5% en poids de B₂O₃, de 10 à 14% en poids de CaO, de 11 à 15% en poids de SrO et de 0 à 2% en poids de ZnO, et essentiellement exempt de MgO, de BaO et d'oxydes de métaux alcalins.